ERRATA

BIOCHIMICA ET BIOPHYSICA ACTA, VOL. 401 (1975)

p. 542, Fig. 1B; abscissa: Numbers 5, 10, 15 and 20 should be replaced by 2.5, 5, 7.5 and 10.

p. 542, legend of Fig. 1, 7th line: 6.2 M^{-1} should read 3.1 M^{-1}.

p. 542, Fig. 2: The experiment is not a "mole fraction experiment". A 2 M sodium-acetate stock solution was mixed on a volume-% basis with a 1 M thallium-acetate solution.

BIOCHIMICA ET BIOPHYSICA ACTA, VOL. 466 (1977) pp. 402–410

In our recent paper presenting calorimetric data on the transformation of a dimyristoyl phosphatidylcholine system from small vesicles to extended bilayers we argued that mechanisms based on diffusive transfer of individual lipid molecules could be discounted because such mechanisms were incapable of accomplishing significant transfer in the short time periods of observation (a few minutes) [1]. The argument was based on a solution adaptation of collision theory which required a crude estimate of average linear velocity of molecules in solution. On further examination we have found that more widely accepted theories of diffusion limited rates do predict transfer to occur within short time periods [2]. Although we feel that the diffusion mechanism is suspect for other reasons, it is important to point out that we should not have argued against it on the basis of absolute rates. Results presented in the original paper are thermodynamic in nature and are, in fact, quite independent of assumed mechanism.

References

1 Kantor, H., Mabrey, S., Prestegard, J. and Sturtevant, J. (1977) Biochim. Biophys. Acta 466, 402–410
2 Weston, R. and Schwarz, H. (1972) Chemical Kinetics, p. 158, Prentice-Hall, Englewood Cliffs, New Jersey

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